

Silver grove school
Class -Viii
Chemistry

Instructions:

The answers to the questions must be written on the paper provided separately.

You will not be allowed to write in the first 15 minutes.

This question paper consists of Section A and Section B.

Attempt all questions from Section A. Section A consists of 15 MCQ type questions.

Attempt all questions from Section B.

Marks for the question or parts are given against each question.

SECTION – A

Question 1 choose the correct option from the following. $(1 \times 15 = 15)$

I. What is homogeneous?

(a).true solution. (b). colloidal. (c). Suspension

(d). Non of these

II. Which one of the following is not a true solution

(a). Sugar solution (b) sand in water. (c). Alcohol in water. (d). Salt solution.

III. The loss of dissolved oxygen from water in bodies is called

a) Oxidation b) Oxygenation c) Deoxygenation d) Eutrophication

IV. Pure water is an/a

a) Acid b) Amphoteric oxide c) Basic oxide d) Neutral

V. Which is used as a catalyst in the harbour bosch process.

(a). Nickel. (b).lead. (c)tin. (d). Silicon

VI. Temporary hardness of water is due to the presence of

a) Magnesium carbonate b) Calcium chloride

c) Magnesium sulphate. d) Calcium sulphate

VII. How many allotropes does carbon have?

a) 1 b) 2 c) 3 d) 4

VIII. Which of the following is a good conductor of electricity?

a) Coal. b) Coke. c) Graphite. d) Diamond

IX. When dilute hydrochloric acid react with sodium carbonate the gas liberated is

(a). Hydrogen (b). Carbon dioxide (c). Methane

(d). Oxygen.

X. The smelling agent added to detect leakage of LPG is

(a). Coal gas. (b). Fuel oil. (c). Ethyl mercaptan. (d). Methane.

XI. Which of the following is not used to put out fire

(a). Water (b). Sand (c). Carbon dioxide (d). Paraffin wax.

Assertion & Reason

a) Both Assertion and Reason are true and Reason is the correct explanation of the Assertion.

b) Both Assertion and Reason are true but Reason is not the correct explanation of the Assertion.

c) Assertion is true but Reason is false.

d) Assertion is false but Reason is true.

XII. Assertion: hydrogen gas can be obtained from water by its electrolysis.

Reason: electrolysis is the process of passing direct electric current through a substance in liquid state.

X III. Assertion: Water hardness is determined by the concentration of calcium and magnesium ions in water.

Reason: Water hardness causes scale build-up in pipes and appliances.

XIV. Assertion: diamond is the purest crystalline form of carbon

Reason :- in a diamond each carbon atom is joined to four other carbon atoms by strong bond

XV. Pure water has a pH value of

- a) 5. b) 6. c) 7. d) 8

Question 2. Complete the blank by using appropriate answer (2)

i.water is saline

ii. Potassium isreactive than calcium

iii. A molecule of hydrogen is represented as.....

iv. Zinc isreactive than iron.

Question 3. Match the following (2)

- | | |
|----------------------|-----------------------------------|
| i). sodium hydroxide | a).H ₃ PO ₄ |
| ii).acetic acid. | b).H ₂ O |
| iii). Water. | c). CH ₃ COOH |
| iv).phosphoric acid. | d). NaOH |

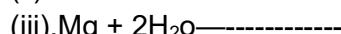
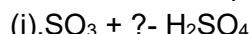
Question 4. Find the odd one out. (3)

i) Diamond, Graphite, Charcoal, Carbon dioxide

ii) Boiling, Freezing, Evaporation, Combustion

iii) Rainwater, River water, Distilled water, Sea water

Question 5. Complete the following reaction. (3)



Section - B

Question 6. Answer in brief (any five). (1×5=5)

(i). Name various amorphous forms of carbon.

(ii). Write any two Chemical properties of hydrogen.

(iii).List any three uses of Hydrogen

(iv). Define the term Allotropy.

(v). What are the Advantages of hard water?

(vi). Why is hydrogen gas use in space rockets

(vii). Name the Crystalline and no crystalline allotropes of carbon .

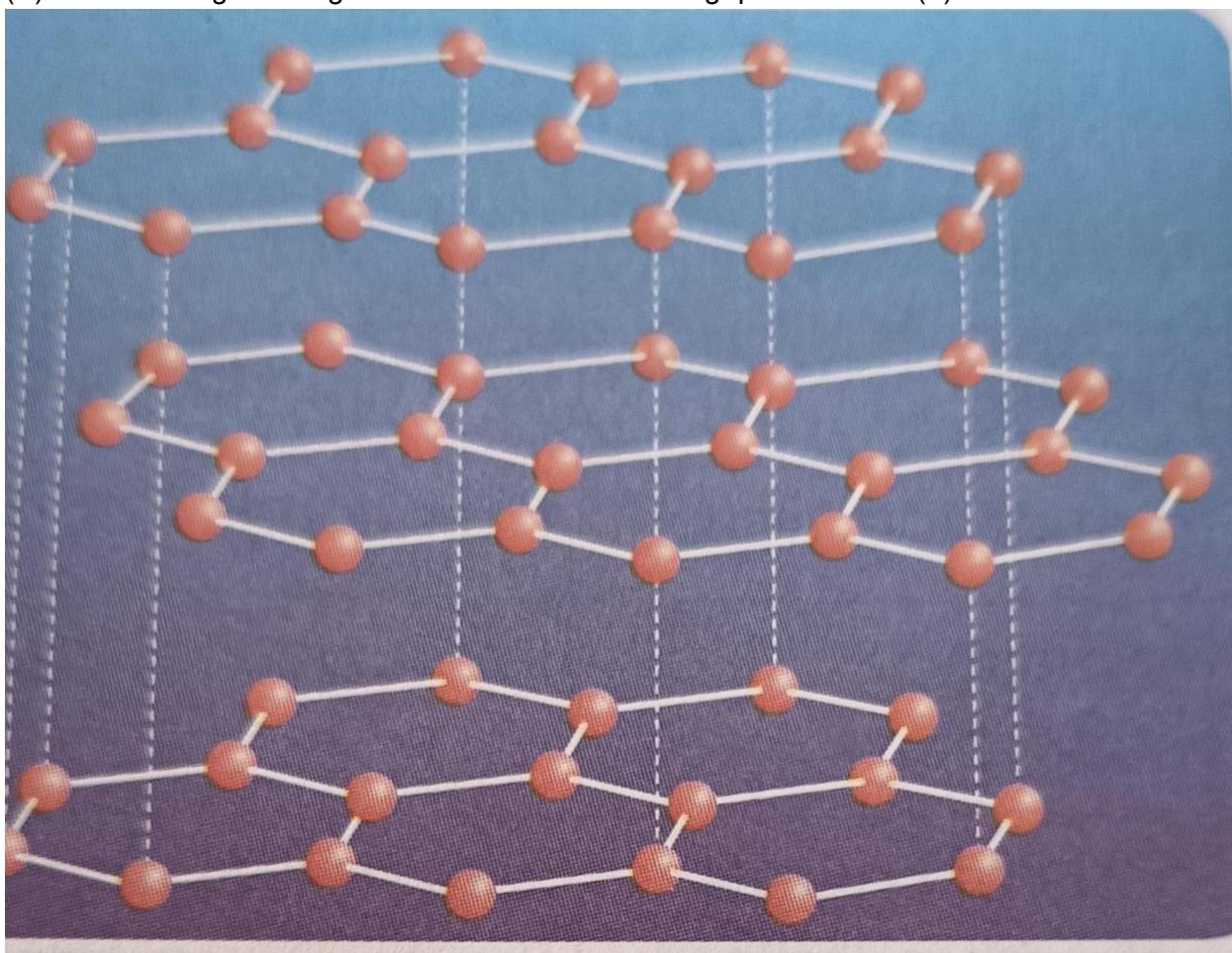
Question 7(A):- Give reason for the following (1×3=3)

(i).Diamond does not conduct electricity whereas graphite does.

(ii). Water is called a universal solvent.

(iii). Hydrogen is added to vegetable oil.

(B). Look at the given diagram and answer the following questions. (2)



(i) Identify the given structure.

(ii). Write the two use of given allotropes of carbon.

**** Project work. (5)**